



Al Sharq Bright International School  
Model Paper for Final Exam 2017-2018

Name: \_\_\_\_\_ Subject: Chemistry Class: 10A Date: \_\_\_\_\_

**Paper – 2**

**A. Multiple choice questions:**

1. What do the  ${}_1\text{H}^1$  hydrogen atoms contain?  
a) electrons and neutrons    b) electrons and protons    c) neutrons only    d) protons only
2. Which property is not characteristic of a base?  
a) It reacts with a carbonate to form carbon dioxide.  
b) It reacts with an acid to form a salt.  
c) It reacts with an ammonium salt to form ammonia.  
d) It turns universal indicator paper blue.
3. What is the relative formula mass,  $M_r$ , of  $\text{CaCO}_3$ ?  
a) 50            b) 68            c) 100            d) 204
4. Hydrochloric acid is used to clean metals. The acid reacts with the oxide layer on the surface of the metal, forming a salt and water. Which word describes the metal oxide?  
a) alloy        b) base        c) element        d) indicator
5. A molecule of X contains two carbon atoms, four hydrogen atoms and two oxygen atoms.  
What is the formula of X?  
a)  $\text{CH}_3\text{COOH}$     b)  $\text{CH}_3\text{CHOOH}$     c)  $\text{CH}_3\text{CHO}$         d)  $\text{CH}_2\text{COOH}$

**Paper – 4**

**B. Answer the Following:**

1. Define the term isotope. Explain Why two isotopes of bromine have the same chemical property.
2. Aluminium is an amphoteric oxide. It is insoluble in water. Describe an experiment to show that aluminium oxide is amphoteric.

3. Silicon (4) oxide has a giant structure.
  - a) Name the type of bonding in silicon (4) oxide.
  - b) Give two physical properties of silicon (4) oxide.
4. Argon is an unreactive noble gas.
  - a) Explain why Argon is unreactive.
  - b) Give one use of Argon.

### Paper – 6

#### C. Alternative to Practical:

Solid E was analysed. E was aluminium salt.

The test on the solid and some of the observation are in the table.

Complete the observation in the table.

Tests	Observations	
Test on solid E a) Appearance of solid E	White crystalline solid.	
b) A little of solid E was heated in a test tube.	Colourless drop of liquid formed at the top of the tube.	
c) A little of solid E was dissolved in distilled water. The solution was dissolved into four test tubes and the following tests were carried out.  1. To the first test tube of solution, drops of aqueous sodium hydroxide was added.	 <hr/> <hr/>	

Excess sodium hydroxide was then added to the test tube.

2. Test 1 was repeated using aqueous ammonia solution instead of sodium hydroxide.

3. To the third test tube of solution, dilute hydrochloric acid was added, followed by barium chloride solution

4. To the fourth test tube of solution, aqueous sodium hydroxide and aluminium powder were added. The mixture was heated.

No reaction

effervescence  
pungent smell given off  
turned damp litmus paper blue

d) What does test b tells you about E

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e) Identify the gas given off in test c 4

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f) What conclusion can you draw about solid E.

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**Note: This is just a model, not the exam paper.**